

Atoms and Molecules

Multiple Choice Questions

Question 1.

Who coined the term 'Parmanu' for the smallest indivisible particles?

- (a) Leucippus
- (b) L. Lavoisier
- (c) Kanad
- (d) None of them

▼ [Answer](#)

Answer: (c) Kanad

Question 2.

The radius of a hydrogen atom is:

- (a) 10^{-2}
- (b) 10^{-4}
- (c) 10^{-9}
- (d) 10^{-10}

▼ [Answer](#)

Answer: (d) 10^{-10}

Question 3.

The most remarkable concept that Dalton's atomic theory proposed was that of the:

- (a) atomic weight
- (b) atomic mass
- (c) molar mass
- (d) none of them

▼ [Answer](#)

Answer: (b) atomic mass

Question 4.

The atomic mass of carbon is:

- (a) 12
- (b) 14
- (c) 16
- (d) 23

▼ [Answer](#)

Answer: (a) 12

Question 5.

Compounds composed of metals and non-metals contain:

- (a) non-charged species
- (b) charged species
- (c) both of them
- (d) none of them

▼ [Answer](#)

Answer: (b) charged species

Question 6.

The formula of ammonium sulphate is:

- (a) NHSO_4
- (b) NH_4SO_2
- (c) NHSO_3
- (d) $(\text{NH}_4)_2\text{SO}_4$

▼ [Answer](#)

Answer: (d) $(\text{NH}_4)_2\text{SO}_4$

Question 7.

The word "mole" was introduced around 1896 by:

- (a) Wilhelm Ostwald
- (b) John Dalton
- (c) Avogadro
- (d) Vergilious

▼ [Answer](#)

Answer: (a) Wilhem Ostwald

Question 8.

The chemical symbol for sodium is:

- (a) So
- (b) Sd
- (c) NA
- (d) Na

▼ [Answer](#)

Answer: (d) Na

Question 9.

Which of the following has a maximum number of atoms?

- (a) 18 g of H_2O
- (b) 18 g of O_2
- (c) 18 g of CO_2
- (d) 18 g of CH_4

▼ [Answer](#)

Answer: (d) 18 g of CH_4

Question 10.

Which of the following contains a maximum number of molecules?

- (a) 1g CO_2
- (b) 1g N_2
- (c) 1g H_2
- (d) 1g CH_4

▼ [Answer](#)

Answer: (c) 1g H_2

[Fill in the Blanks.](#)

Question 11.

In a compound such as water, the ratio of the mass of hydrogen to the mass of oxygen is always _____

▼ [Answer](#)

Answer: 1 : 8

Question 12.

Atomic radius is measured in _____

▼ [Answer](#)

Answer: nanometre

Question 13.

The symbol of iron is Fe from its Latin name _____

▼ [Answer](#)

Answer: Ferrum

Question 14.

One atomic mass unit is a mass unit equal to exactly _____ the mass of one atom of carbon-12.



▼ [Answer](#)

Answer: th

Question 15.

The formula unit mass of a substance is a sum of the atomic _____ of all its constituent atoms.

▼ [Answer](#)

Answer: masses

Question 16.

Avogadro number is represented by _____

▼ [Answer](#)

Answer: N_0

Question 17.

The relative mass of atoms of all elements is obtained by comparing them with the mass of a _____ atom.

▼ [Answer](#)

Answer: Carbon-12

Question 18.

The Avogadro constant 6.022×10^{23} is defined as the number of atoms in exactly _____ of carbon-12.

▼ [Answer](#)

Answer: 12 g

[True/False.](#)

Question 19.

An atom is the smallest particle of the element that cannot usually exist independently and retains all its chemical properties.

▼ [Answer](#)

Answer: True

Question 20.

The Law of conservation of mass states that mass can neither be created nor destroyed in

a chemical reaction.

▼ [Answer](#)

Answer: True

Question 21.

The mass of 1 mole of a substance is equal to its relative atomic or molecular mass in grams.

▼ [Answer](#)

Answer: True

Question 22.

16u oxygen has 16 atoms of oxygen.

▼ [Answer](#)

Answer: False

Question 23.

The atomic mass of calcium is 40.

▼ [Answer](#)

Answer: True

Question 24.

Aluminium is diatomic.

▼ [Answer](#)

Answer: False

[Match the Column.](#)

Question 25.

A	B
1. 1 mole of carbon atoms	(i) 1 g of H atoms
2. 1 mole of molecules	(ii) Molecular mass in gram
3. 1 mole of hydrogen atoms	(iii) Relative mass of those particles in gram

4. 1 mole of any particle (atoms, molecules, ions)	(iv) 12-gram carbon atoms
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▼ [Answer](#)

Answer:

A	B
1. 1 mole of carbon atoms	(iv) 12-gram carbon atoms
2. 1 mole of molecules	(ii) Molecular mass in gram
3. 1 mole of hydrogen atoms	(i) 1 g of H atoms
4. 1 mole of any particle (atoms, molecules, ions)	(iii) Relative mass of those particles in gram

[Answer in one Word/Sentence.](#)

Question 26.

What is the value of the Avogadro number?

▼ [Answer](#)

Answer: 6.022×10^{23}

Question 27.

How many atoms of hydrogen are there in 1g of hydrogen?

▼ [Answer](#)

Answer: 6.022×10^{23}

Question 28.

What name is given to a group of atoms carrying a charge?

▼ [Answer](#)

Answer: Polyatomic ions

Question 29.

What term is given to the mass of one-mole atoms of a substance?

▼ [Answer](#)



Answer: Molar mass

Question 30.

Write the chemical formula of calcium hydroxide.

▼ [Answer](#)

Answer: Ca(OH)_2

Question 31.

Write the chemical formula of sulphuric acid?

▼ [Answer](#)

Answer: H_2SO_4

